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# Stoichiometry

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Chapter 12 Stoichiometry

Practice Problems Chapter 12

Stoichiometry Practice

Problems Answer Key A In any

stoichiometry problem, the

first step is always to

calculate the number of

moles of each reactant

present. In this case, we

are given the mass of  $K_2Cr$



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Key 2.07 in 1 mL of solution,  
which we can use to  
calculate the number of  
moles of  $K_2Cr$  ...

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 $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  a. If you use 2.3 moles of Al, how many moles of  $\text{Al}_2\text{O}_3$  can you make? b. If you want 3.9 moles of  $\text{Al}_2\text{O}_3$ , how many moles of  $\text{O}_2$  are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate:  $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$  ...

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Terms in this set (...)

Stoichiometry. that portion of chemistry dealing with the numerical relationships in chemical reactions . What is stoichiometry based on? the law of conservation of mass. What does stoichiometry involve? balancing chemical equations and mole ratios. mole ratio.

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**Key** a conversion factor that relates the number of moles of any two ...

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